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UNIT 3 HOMEWORK 1 SOLUTIONS

Unit 3 Homework 2 Solutions Total 46 marks REFRACTION 200017 C 118 C 127 c) i) 42° 1 ii) $n = 1 \sin \theta_c = 1 \sin 42^\circ = 1.49$ (1) 2 iii) Decreases to 0 Wm⁻² 1 6200117 C 127 a) $n = d \sin \theta = 2 \times 486 \times 10^{-9} = 2.16 \times 10^{-6}$ x $\sin \theta = 26.7^\circ$ (1) 2 b) i) $n = \sin \theta_a \sin \theta_m = \sin 47^\circ \sin 27^\circ = 1.61$ (1) 2 ii) $n = 1 \sin \theta_c = 1.61 = 1 \sin \theta_c$ $\theta_c = 38.4^\circ$ (1) 2 The ray will be totally internally ref...

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TEAMS: BAT ENG. DESIGN UNIT 4.4A HOMEWORK ANSWERS

080226 Center of Mass Problems E Find the CM relative to the front of the car. *zfron, xfron, *ffibu kxtu k - LM - ffi*x* m + m w trmt + m. Dack (roso rg)(2.s0 m) + z(tl.0 rcg)(z.ro m) + (zo.o rg)(: .eo m) = di4A2. The point that will follow a parabolic trajectory is the center of mass. Find the CM relative to the bottom of the mallet. Each part of the hammer (handle and head) can be treated as a point mass...

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MICROSOFT WORD - CH. 12 SECTION 12.6 TEXTBOOK HOMEWORK ANSWERS

AP Chemistry Chapter 12 Answers Zumdahl 12.49 For elementary reactions, the rate law can be written using the coefficients in the balanced equation to determine orders. a. Rate = $k[\text{CH}_3\text{NC}]$ b. Rate = $k[\text{O}_3][\text{NO}]$ c. Rate = $k[\text{O}_3]$ d. Rate = $k[\text{O}_3][\text{O}]$ 12.50 The observed rate law for this reaction is: Rate = $k[\text{NO}]^2[\text{H}_2]$. For a mechanism to be plausible, the sum of all the steps must give the overall balanced equation...

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MICROSOFT WORD - CH. 14 SECTIONS 14.1-14.2 TEXTBOOK HOMEWORK ANSWERS

AP Chemistry Chapter 14 Answers Zumdahl 14.27 a. $\text{HClO}_4(\text{aq}) + \text{H}_2\text{O}(\text{l}) \rightleftharpoons \text{H}_3\text{O}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. Only the forward reaction is indicated since HClO_4 is a strong acid and is basically 100% dissociated in water. For acids, the dissociation reaction is commonly written without water as a reactant. The common abbreviation for this reaction is: $\text{HClO}_4(\text{aq}) \rightleftharpoons \text{H}^+(\text{aq}) + \text{ClO}_4^-(\text{aq})$. This reaction is also called the K_a re...





















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